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The crystal structure determinations of transdicyano trien cobalt (III) perchlorate and potassium antimony (III) pentachloride and two phase determination techniques based on Patterson deconvolution

Robert Kingsley Wismer *Iowa State University*

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**Iowa State University, Ph.D., 1972 Chemistry, physical** 

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The crystal structure determinations of trans-dicyano trien cobalt(III) perchlorate and potassium antimony(III) pentachloride

and

two phase determination techniques based on Patterson deconvolution

by

Robert Kingsley Wismer

A Dissertation Submitted to the Graduate Faculty in Partial Fulfillment of The Requirements for the Degree of DOCTOR OF PHILOSOPHY

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Major: Physical Chemistry

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#### INTRODUCTION

This thesis describes the crystal structures of two compounds and details two methods for the determination of the phases of structure factors. The crystal structures were undertaken to extend the currently available knowledge of cobalt and antimony complexes.

The cobalt compound was believed by its discoverers to be cis-a(NC)<sub>2</sub> trien cobalt(III) perchlorate. Thus, it would have been the sole known mononuclear cobalt complex in which the cyanide group was coordinated to cobalt through the nitrogen atom. The structure that was determined is actually trans- $\text{(CN)}_{2}$  trien cobalt(III) perchlorate and is one of the few examples in which a quadradentate ligand, not restricted by its own structure to be planar, is equatorially coordinated to cobalt.

The second structure described is that of  $K_{\mathcal{D}}SbCl_{\mathsf{G}}$ . The structure of the  $SbCI_5^-$  anion has only been determined once previously, in the compound  $(NH^4)$ <sub>2</sub>SbCl<sub>5</sub>. In that instance, the disagreement between observed and calculated structure factors was such that the uncertainty in bond distances was  $0 \cdot 5 \, 0.6$  Structures of antimony halides are known to be very sensitive to the cation present.

The first of the phase determining methods applies the techniques of pseudo heavy atom superposition and then uses space group symmetry to determine atomic

locations in electron density space. All of these operations are performed by a computer program with a minimum of human intervention.

The second phase determining method also uses superposition techniques, to enhance the phase determining properties of a pseudo electron density map which has been computed from origin-defining reflections. The coefficients of the Fourier transform of this modified map are then used to phase those reflections with the largest normalized unitary structure factors. This method greatly simplifies one of the major problems of direct methods, the expansion of the origin-defining set.

## THE CRYSTAL AND MOLECULAR STRUCTURE OF TRANS-DICYANO TRIETHYLENETETRAMIME COBALT(III) PERCHLORATE

### Introduction

The bonding mode of cyanide to cobalt has been the subject of considerable controversy in recent years. The energy difference between the cyano and isocyano coordinations is small enough to permit either to occur.<sup>1</sup> However, neutron diffraction studies of cyano-cobalt complexes<sup>2</sup>,<sup>3</sup>,<sup>4</sup> have regularly shown that the cyanide group bonds through carbon. Since it is ofttimes difficult to distinguish between atoms which differ by but one electron by X-ray diffraction, bonding via the carbon has sometimes been assumed.<sup>5</sup>

Kuroda and Gentile<sup>G</sup> recently have prepared what they characterize as an isocyano cobalt complex,  $cis-\alpha-isocyano$ triethylenetetramine cobalt(III) perchlorate. Their assignment of the isocyano coordination is based on visible and ultraviolet spectroscopy. Because there is some doubt as to the correctness of this assignment, it was decided to carry out an X-ray structure determination of this compound.

#### Experimental

A sample of the compound, prepared by the method of Kuroda and Gentile,  $6$  was kindly supplied by Dr. James H.

Espenson. Amber-colored crystals were obtained by recrystallization from an aqueous solution. High resolution infrared spectra of the initial sample and of the recrystallized material, both in KBr pellets, were virtually identical, indicating that isomerization had not occurred during solution or recrystallization. Microscopic examination revealed that the crystals have sharply defined faces and are needle-like in appearance with a distorted hexagonal cross section in which alternate sides are of unequal lengths.

Crystals were selected and mounted on a glass fiber with Duco cement thinned with amyl acetate. Preliminary Weissenberg and precession photographs exhibited 2/m Laue symmetry, indicating a monoclinic space group. The following systematic absences were observed:  $h0\ell$  when h +  $\ell$  = 2n + 1, and OkO when  $k$  = 2n + 1. These absences are only consistent with the space group  $P2^1/n$ . The unit cell parameters at 25° C are  $a = 9.8544 + 0.0015$ , b = 22.3529 ± 0.0045, Ç = 6.6766 + 0.0020 **1,** and  $\beta$  = 100.85 + 0.03°. These parameters and their standard  $\mathsf{R}$ deviations were obtained by a least squares fit<sup>o</sup> to the two-theta values of twelve independent reflections whose centers were determined by top-bottom, left-right beam splitting on a previously aligned Hilger-Watts four-circle diffractometer (MoKa radiation,  $\lambda = 0.71069 \, \AA$ ). Any

4-

error in the instrumental zero was eliminated by centering the reflection at both plus two-theta and minus two-theta. A calculated density of 1.52 g/cc for four molecules per unit cell agrees quite well with an observed density of 1.65 + 0.01 g/cc, which was determined by flotation techniques.

For data collection, a crystal was selected having approximate dimensions  $0.1 \times 0.1 \times 0.9$  mm along the a, b, and c crystal axes, respectively, and was mounted such that the c axis coincided with the  $\beta$  axis of the diffractometer. Data were collected at room temperature using a Hilger-Vatts four-circle diffractometer interfaced to an SDS 910 computer in a real time mode, equipped with a scintillation counter, and using Zr-filtered MoKa radiation. Within a two-theta sphere of 60° (sin $\theta/\lambda = 0.704 \text{ }\text{\AA}^{-1}$ ) all data in the hk<sup> $\ell$ </sup> and  $\bar{h}$ k $\ell$  octants were recorded using the  $\theta$ -20 scan technique with a take-off angle of  $5.5^\circ$ . Symmetric scan ranges of 1.2° in 20 at low two-theta values to 2.0° at large two-theta values were used. Stationarycrystal, stationary-counter background counts of half the scan time were taken at the beginning and end of each scan. A counting rate of 0.4096 seconds per step of 0.01° in  $\theta$ was employed. A total of 5081 reflections were measured in this way.

As a general check on electronic and crystal stability, the intensities of three standard reflections were remeasured periodically during the data collection period. These reflections did not vary to any significant degree during the entire period of data collection.

Based on a linear absorption coefficient of  $\mu$  = 14.982 cm<sup>-1</sup>, the maximum and minimum transmission factors<sup>9</sup> were 87.86% and 85.96% respectively, and no absorption correction was deemed necessary.

The intensity data were corrected for Lorentzpolarization effects. The estimated error in each intensity was calculated by

$$
\sigma_{\rm I}^2 = C_{\rm T} + C_{\rm B} + (0.03 C_{\rm T})^2 + (0.05 C_{\rm B})^2
$$

where  $C_{\phi}$  and  $C_{\phi}$  are the total count and the background count, respectively. The factor 0.05 represents an estimate of non-statistical errors. The estimated deviations in the structure factors were calculated by the finite difference method. $^{10}$  Of the 4678 independent reflections, 3314 were considered observed (>  $3\sigma_{\overline{F}}$ ).

#### Solution and Refinement

The positions of the cobalt and chlorine atoms were obtained from analysis of a sharpened Patterson function.<sup>11</sup> The remaining non-liydrogen atoms were found by successive structure factor  $12$  and electron density map calcula-

 $tions.$  13,14 These atomic positions were then refined by a full matrix least squares procedure,<sup>12</sup> minimizing the function  $\sum w (|F_o| - |F_c|)^2$ , where  $w = 1/\sigma_F^2$ , to a conventional discrepancy factor of R =  $\Sigma$ || $F^{\text{}}_{\text{O}}$ | -| $F^{\text{}}_{\text{C}}$ || $\Lambda$ | $F^{\text{}}_{\text{O}}$  = 0.170. The scattering factors used were those of Hanson  $\underline{\text{et al.}}^{15}$  with the cobalt and chlorine scattering factors modified for the real and imaginary parts of anomalous dispersion.  $^{16}$ 

An electron density difference map verified that all the non-hydrogen atoms had been accounted for, but indicated that some anisotropic motion, particularly of the cobalt and chlorine atoms, was present. Accordingly, anisotropic refinement was begun and after six cycles of refinement, values of R and R (=  $\sum_{\omega}(|F_{o}|-|F_{c}|)^{2}/\omega|F_{o}|^{2})^{1/2}$ ) of 0.154 and 0.169, respectively, were obtained. The positions of the hydrogens bonded to carbons in the triethylenetetramine ligand were calculated with an H-C-C angle of 109-5° and O a C-H distance of I**.07** A. The hydrogen atoms were assigned isotropic thermal parameters of  $3.0 \text{ }\overset{\circ}{A}^2$ . Two cycles of refinement with anisotropic thermal parameters for all nonhydrogen atoms, varying only the positional and thermal parameters of the non-hydrogen atoms and the overall scale factor, followed by a recalculation of the hydrogen atom positions and two further cycles of least squares refinement, yielded values for R and  $w R$  of 0.123

and 0.156, respectively.

A final electron density difference map showed no  $2,3$ peaks greater than  $1.2 e/\tilde{A}^2$ . A final statistical analysis of the  $F_o$  and  $F_c$  values as a function of the scattering angle and magnitude of  $F_{\alpha}$  revealed no unusual trends and suggests that the relative weighting scheme used is a reasonable one.

The bonding orientation of the two cyanide groups was determined after the final least squares refinement. The atom multipliers of the two cyanide group atoms were allowed to vary under two sets of conditions. In the first of these, carbon scattering factors were assigned to  $C(7)$ and 0(8) and nitrogen scattering factors were assigned to  $N(5)$  and  $N(6)$ . In the second instance, the scattering factor assignments were reversed. All four atom multipliers were initially set at a value of 1.00. The results of these refinements are shown in Table I and clearly indicate the correct assignments.

The final positional and thermal parameters for the non-hydrogen atoms are listed in Table II. The standard deviations were calculated from the inverse matrix of the final least squares refinement cycle. The calculated hydrogen atom positions are listed in Table III. Bond

	Assigned Scattering	Atom Multiplier Theoretical Actual				
Atom	Factor Table		$Co-C-N$	$Co-N-C$	R	ωR
C(7)	Carbon	1.033	1.0	1.17		
C(8)	Carbon	1.034	1.0	1.17		
N(5)	Nitrogen	1.026	1.0	0.86		
N(6)	Nitrogen	1.046	1.0	0.86		12.4 15.8
C(7)	Nitrogen	0.823	0.86	1.0		
C(8)	Nitrogen	0.825	0.86	1.0		
N(5)	Carbon	1.293	1.17	1.0		
N(6)	Carbon	1.314	1.17	1.0		12.4 15.8

Table I. Determination of the bonding mode of cyanide in  $trans-(CN)$ <sub>2</sub> trien  $Co^{+++}ClO_A$ 

 $\mathcal{L}_{\text{max}}$ 



**Table II Final positional and anisotropic thermal parameters of non-hydrogen atoms in trans-(NC)<sub>2</sub>-triem Co<sup>III</sup>ClO<sub>4</sub> a, b** 

**® In this and subsequent tables, numbers In parentheses represent standard deviations in the least significant digits.** 

**b** Anisotropic thermal parameters are defined by:  $T = exp(-({h^2\rho_{11} + \kappa^2\rho_{22} + t^2\rho_{33} + 2h\kappa\rho_{12} + 2\kappa\ell\rho_{23} + 2h\ell\rho_{13})}).$ 



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Table III. Calculated hydrogen atomic positions in trans- $(CN)$ <sub>2</sub> trien Co<sup>111</sup>ClO<sub>4</sub>

lengths and bond angles, and significant non-bonded distances are listed in Tables IV and V, respectively, along with their standard deviations.  $17$  The final values of the observed and calculated structure factors are listed in Table VI.

Shortly after the data had been collected and the crystal removed from the diffractometer, the instrument was found to be slightly misaligned. Thus, the data are somewhat poorer than might be expected. However, since all the stereochemical features of interest were well determined and the electron density difference map was rather featureless, the data were not retaken.

#### Description of the Structure

The results clearly indicate that the bonding of cyanide to the cobalt is through the carbon and not the nitrogen, and that these groups are trans and not cis to one another, contrary to the predictions of Kuroda and Gentile.<sup>6</sup> Hence, the compound is trans-dicyanotriethylenetetramine cobalt(III) perchlorate, and is illustrated in Figure 1.<sup>18</sup> This configuration has been anticipated by Konya, Nishikawa, and Shibita<sup>7</sup> on the basis of spectroscopic examination of similar compounds. The trans configuration is also consistent with an exceedingly sharp band for the C-N stretching frequency at 2145  $cm^{-1}$ , with no evidence of splitting.

			4
$C1 - O(1)$	1.427(9)	$C(8)-N(6)$	1.180(13)
$C1 - O(2)$	1.424(8)	$N(1)-C(1)$	1.462(15)
$C1 - O(3)$	1.456(9)	$N(2)-C(2)$	1.488(12)
$C1 - O(4)$	1.392(10)	$N(2)-C(3)$	1.476(13)
$Co-N(1)$	1.978(8)	$N(3)-C(4)$	1.484(14)
$Co-N(2)$	1.951(8)	$N(3)-C(5)$	1.444(15)
$Co-N(3)$	1.922(8)	$N(4)-C(6)$	1.506(15)
$Co-N(4)$	2.001(8)	$C(1)-C(2)$	1.527(16)
$Co-C(7)$	1.909(12)	$C(3)-C(4)$	1.551(17)
$Co-C(8)$	1.894(12)	$C(5)-C(6)$	1.536(18)
$C(7)-N(5)$	1.161(14)		
$O(1)-Cl-O(2)$	110.7(6)	$N(4) - Co - C(8)$	91.5(4)
$O(1)-Cl-O(3)$	109.1(6)	$C(7)-Co-C(8)$	177.9(4)
$O(1)-Cl-O(4)$	109.3(6)	$Co-N(1)-C(1)$	106.7(6)
$O(2)-Cl-O(3)$	107.1(6)	$Co-N(2)-C(2)$	108.1(7)
$O(2)-Cl-O(4)$	112.3(7)	$Co-N(2)-C(3)$	108.9(7)
$O(3)-Cl-O(4)$	108.4(6)	$Co-N(3)-C(4)$	110.8(7)
$N(1) - Co-N(2)$	88.0(4)	$Co-N(3)-C(5)$	109.0(7)
$N(1)-CO-N(3)$	172.5(4)	$Co-N(4)-C(6)$	106.3(7)
$N(1) - Co-N(4)$	99.4(4)	$Co-C(7)-N(5)$	173.7(10)
$N(1) - Co - C(7)$	89.2(4)	$Co-C(8)-N(6)$	174.1(9)

Table IV. Selected bond distances and angles in angstroms and degrees for cis- $(\text{CN})_{2}$  trien  $\text{Co}^{+++} \text{ClO}_{4}$ 

Table IV (Continued)

$N(1)-Co-C(8)$	89,2(4)	$N(1)-C(1)-C(2)$	111.3(9)
$N(2)-Co-N(3)$	84.5(4)	$N(2)-C(2)-C(1)$	105.9(8)
$N(2)-Co-N(4)$	172.6(4)	$N(2)-C(3)-C(4)$	111.4(8)
$N(2)-Co-C(7)$	91.8(4)	$N(3)-C(4)-C(3)$	110.2(9)
$N(2)-Co-C(8)$	89.6(4)	$N(3)-C(5)-C(6)$	108.3(10)
$N(3)-Co-N(4)$	88.1(4)	$N(4)-C(6)-C(5)$	109.1(9)
$N(3)-Co-C(7)$	90.7(4)	$C(2)-N(2)-C(3)$	116.8(8)
$N(3)-Co-C(8)$	91.0(4)	$C(4)-N(3)-C(5)$	114.5(10)
$N(4)-Co-C(7)$	87.3(4)		



 $a$  Primed atoms indicate atoms in another molecule.

Table V. Significant non-bonded distances in  $trans-(CN)_{2}$ 



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Table VI. Observed and calculated structure factors

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Figure 1. A formula unit of trans-cyano triethylenetetramine cobalt(III) perchlorate

The cyanide groups are within a degree of being perpendicular to the nitrogens of the trien ligand (Table IV), and the variations from an ideal octahedral configuration displayed by the K-Co-K angles appear to simply reflect the restrictions imposed by the quadradentate ligand (Figure 2). This trien ligand coordinated to cobalt in a cis configuration has been investigated by Freemen and Maxwell,  $^{19}$  and Dwyler and Maxwell;  $20$  their results have been used as the basis of many of the comparisons which follow.

The  $N(1)$ -Co- $N(2)$  and  $N(3)$ -Co- $N(4)$  angles of 88.0 and 88.1° are reasonable and consistent with similar results obtained in other investigations of the trien ligand. A slight compression of the  $Co-N(2)-C(3)-C(4)-N(3)$  ring due to the equatorial coordination of the ligand adequately explains the somewhat small value of 84.5° for the  $N(2)-Co-N(5)$  angle. Hydrogon-hydrogen repulsions would be expected to produce an increase in the  $N(L)-CO-N(4)$  angle compared to the nominal 90°, and the 99° angle found is quite reasonable. À similar result was obtained during the investigation of a cobalt compound containing a homologous ligand.  $21$  The bond angles within the trien ligand, with the exception of the C-N-C angles, do not differ from the expected tetrahedral angles by more than three standard deviations. In addition, they fall within two standard deviations of the range of values of 108.4°



Figure 2. Bond distances and angles within the trien ligand

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to 112.3° for Co-N-C bonds, and of 105.2° to 111.4° for N-C-C bonds reported by other investigators of this ligand. The C-N-C bond angles of 116.8° and 114.6° are also reasonable in view of the fact that the Co-N-C angles are, on the average, somewhat less than tetrahedral. These C-N-C angles also agree within three standard deviations with other results obtained in investigations of this ligand.

The bond distances within the trien ligand range from 1.92 to 2.00  $\stackrel{\circ}{\Lambda}$  for Co-N bonds, from 1.44 to 1.51  $\stackrel{\circ}{\Lambda}$  for **O**  N-C bonds, and from 1.55 to 1.55 A for C-C bonds, with average values of 1.96, 1.48, and 1.54 A, respectively. These distances in no instance differ significantly from previously reported values for other compounds involving the trien ligand coordinated to cobalt.

The four trien nitrogens and the cobalt atom all fall within 0.01 A of the least squares plane. The three fivemembered rings formed by the quadradentate ligand and the cobalt atom are not planar, however, as the carbon atoms are located above and below the N-Co-N planes. Nonetheless, all of the atoms in the trien ligand fall within 0.37 A of a least squares plane.

As has been described above, the coordination of the cyanide groups to the cobalt is clearly through the carbon. The Co-C average bond distance of 1.90 A is in good

agreement with previously published values. The average C-N distance of 1.17 Å also agrees quite well with the range of previously reported values,  $1.15$  to  $1.18$   $A.$  <sup>2</sup>, <sup>22-24</sup> The fact that no part of the chain  $N(6)-C(8)-Co-C(7)-N(5)$ is linear, with bond angles of  $N(6)-C(8)-Co$ ,  $C(8)-Co-C(7)$ , and  $Co-C(7)-N(5)$  of  $174.1$ ,  $177.9$ , and  $175.7^{\circ}$ , respectively, is somewhat unusual but not surprising. Ourry and Runciman<sup>2</sup> have already observed such a phenomenon. Much of this can be attributed to steric effects both inside the moiety and between moieties (Figure 3). Significant non-bonded distances are given in Table V.

The C1-0 bond distances in the perchlorate group average 1.42 A and the angles average 109.5°. The relatively large thermal parameters, possibly due to spatial as well as temporal disorder probably account for the large individual deviations from the average values. However, these average values agree reasonably well with previously published values for this anion.<sup>21</sup> A weak hydrogen bond is possible between  $N(4)$  and  $O(3)$ , which are 3.03 Å apart. Such a hydrogen bond would help stabilize the perchlorate group.



 $\sim 10^{-1}$ 

 $\Delta$ 

Figure 3. Unit cell showing packing of trans-cyano triethylenetetramine cobalt(III) perchlorate

 $\sim 10^7$ 

# THE STRUCTURE OF POTASSIUM ANTIMONYCIII) PENTACHLORIDE

#### Introduction

Structural investigations of antimony halide complexes have been actively pursued in this laboratory during recent years, beginning with the work of Lawton and Jacobson.<sup>25</sup> The structures of species containing antimony-(III) are of particular interest because of questions concerning the stereochemical role of the lone pair of electrons.

Only a very few antimony (III) chloride structures have been reported in the literature. SbCl<sub>3</sub> exhibits a trigonal pyramid structure with antimony at the apex,  $26$ while in  $(\text{NH}_{\mu})_{2}$ SbCl<sub>5</sub> the chlorine atoms are arranged in a square pyramidal coordination geometry around the antimony.<sup>27</sup> In [C<sub>5</sub>H<sub>5</sub>NH] [Sb<sup>III</sup>Cl<sub>*L*</sub>], the halogen configuration around the antimony can be described as tetragonally distorted octahedral with two of the equitorial atoms  $28<sub>q</sub>$ removed,<sup> $\zeta$ </sup> and SbCl<sub> $\zeta$ </sub><sup>2-</sup> forms perfect octahedra in Co(NH<sub>z</sub>)<sub> $\zeta$ </sub>  $SbCl_{6}$ <sup>29</sup>

Because of the relatively poor determination  $(R = 0.24)$ of the structure of  $(NH<sub>4</sub>)_{2}SbCl<sub>5</sub>$  and because of the seeming variety of structural geometries possible for antimony (HI), it was decided to undertake a structure determination of  $K_2SbCl_5$ .

#### Experimental

A sample of  $K^{\text{}}_{2}$ SbCl<sub>5</sub> was kindly supplied by Dr. Donald Macalady, who prepared the compound in the following manner: Two solutions, one of  $SbCl_{\overline{3}}$  in 3N HCl and one of KCl in 5N HCl, were combined so that the molar ratio of potassium to antimony in the resulting solution was 5:7. This solution was allowed to stand, and crystals formed by slow evaporation. Microscopic examination revealed many different crystal morphologies. Thin plates appeared to be the basic unit of all these morphologies, however.

A crystal was cut from one of these plates and mounted on the end of a glass fiber with Duco cement thinned with amyl acetate. Preliminary Weissenberg and precession photographs exhibited 2/m Laue symmetry, indicating a monoclinic space group. The following systematic absences were observed:  $h0$ <sup>t</sup> when  $l = 2n + 1$ , and OkO when  $k = 2n + 1$ . These absences are only consistent with the space group  $P2<sub>1</sub>/c$ . The unit cell parameters at 25°C are  $\underline{a}$  = 8.8686  $\underline{+}$  0.0007,  $\underline{b}$  = 12.4577  $\underline{+}$  0.0013,  $\underline{c}$  = 8.9280  $\underline{+}$ 0.0013 Å and  $\beta$  = 110.512 + 0.011°. These parameters and their standard deviations were obtained by a least squares  $\mathbf{a}$ fit<sup>8</sup> to the 20 values of sixteen independent reflections whose centers were determined by left-right, top-bottom beam splitting on a previously aligned four-circle diffractometer (MoKa radiation,  $\lambda = 0.71069$  A). A calculated

density of 2.711 g/cc for four molecules per unit cell agrees quite well with an observed density of  $2.72 +$ 0.01 g/cc, determined by flotation techniques.

For data collection the crystal described above, measuring approximately 0.2 mm along each of the crystal axes, was mounted so that the ç axis coincided with the  $\phi$  axis of the diffractometer. Data were collected at room temperature using an automated four-circle diffractometer designed and built in the Ames Laboratory. The upper full circle was purchased from STOE and is equipped with encoders (Baldwin Optical) and drive motors. The design of the base allows the encoders to be directly connected to the main  $\theta$  and  $2\theta$  shafts, using solid and hollow shaft encoders, respectively. The diffractometer is interfaced to a PDP-15 computer in a real time mode and is equipped with a scintillation counter. Zirconiumfiltered MoKa radiation was used for the data collection. A scan rate of 0.1 second per step of  $0.01^\circ$  in  $\theta$  was employed with a variable scan range of 35 steps plus 1 step per degree theta. Stationary-crystal, stationarycounter background counts of a quarter of the scan time were taken at the beginning and end of each scan. Before the scan was made a peak height measurement was used to determine if the reflection was observed. To be scanned, the reflection had to exceed the background by

more than six counts. If the reflection met this criterion, the w setting was then adjusted slightly, if necessary, to maximize the peak height. Within a twotheta sphere of  $50^{\circ}$  (sin $\theta/\lambda = 0.595 \text{ \AA}^{-1}$ ), all data in the hk<sup> $t$ </sup> and  $\bar{h}$ k  $t$  octants were measured in this manner, using a take-off angle of 4.5°. Of the 1750 reflections examined, 1614 met the peak height criterion and were scanned.

As a general check on electronic and crystal stability, the intensities of three standard reflections were remeasured every twenty-five reflections. These reflections did not vary to any significant degree during the entire period of data collection.

The intensity data were corrected for Lorentzpolarization effects and for effects due to absorption. An absorption correction was made using the Tompa-Alcock absorption correction  $program^{30}$ ,  $31$  using a linear absorption coefficient of  $\mu$  = 52.28 cm<sup>-1</sup>. The maximum and minimum transmission factors were 43.95% and 35.98%, respectively. The estimated error in each intensity was calculated by

$$
\sigma_{\rm T}^{2} = c_{\rm T} + 2c_{\rm B} + (0.03 \, c_{\rm T})^{2} + (0.03 \, c_{\rm B})^{2} + (0.03 \, c_{\rm B})^{2}
$$

where  $C^T, C^B$ ,  $C^T$ , and  $T^A$  are the total count, the

background count, the net count, and the transmission factor, respectively, and the factor 0.03 represents an estimate of non-statistical errors. The estimated deviations in the structure factors were calculated by the finite difference method.  $10$  Of the 1614 independent reflections, 1543 were considered observed (> 2.0 $\sigma$ <sub>F</sub>).

#### Solution and Refinement

The position of the antimony atom was obtained from analysis of a sharpened three-dimensional Patterson function. $^{11}$  The remaining atoms were found by successive structure factor<sup>12</sup> and electron density map calculations.<sup>14</sup> These atomic positions were refined by a full matrix least squares procedure,  $^{12}$  minimizing the function  $\sum w (|F_{o}|-|F_{c}|)^{2}$ , where  $\omega = 1/\sigma_{\overline{F}}^2$ , to a conventional discrepancy factor of  $R = 0.169$ . The scattering factors used were those of Hanson  $\underline{et}$  al.,<sup>15</sup> modified for the real and imaginary parts of anomalous dispersion.<sup>16</sup>

An electron density difference map verified that all the atoms had been accounted for, but indicated that some anisotropic motion, particularly of the antimony, was present. Accordingly, anisotropic refinement was begun and after four cycles of refinement, values of R and of 0.054 and 0.078, respectively, were obtained.

A final electron density difference map showed no  $.5^{\circ}$ peaks greater than  $0.9 e/\tilde{\Lambda}^2$ . A final statistical analysis of  $\overline{w\wedge}^2$ , where  $\wedge$  = ( $\left| \begin{array}{c} F \end{array} \right| + F_c$ ), as a function of the scuttering angle and magnitude of  $\mathbb{F}_0$  revealed no unusual trends and suggests tliat the relative weighting scheme used is a reasonable one.

The final positional and thermal parameters are listed in Table VII. The standard deviations were calculated from the inverse matrix of the final least squares cycle. Bond lengths, bond angles, and significant non-bonded distances $^{17}$  are listed in Table VIII. The final values oi" tlio observed and calculatod structure factors are listed in Table IX.

#### Description and Discussion

The configuration of  $SbCl_5^{=}$  in  $K_2SbCl_5$  is essentially that of a square pyramid which has been distorted by interionic and packing forces. The inversion-related anions are packed approximately base to base, as depicted in Figure  $4 \cdot ^{18}$  The axial antimony-chlorine bond length is 2.385 Å, in good agreement with the axial distance in  $(N\mathbb{H}_{4})_{2}$ SbCl<sub>5</sub> (2.36  $\hat{\Lambda}$ ),<sup>27</sup> the shorter distance (2.38  $\hat{\Lambda}$ ) in  $\inf_{\sigma_j\in\sigma_j}C_{\sigma_j}\inf_{\sigma_j}S_{\sigma_j}$  and only slightly longer than the distance  $(2.32 \text{ Å})$  in SbCl<sub>3</sub>.<sup>26</sup>


Table 7II. Final atonic coordinates and thermal parameters (x  $10^{4}$ ) for  $K_pSbOL_g^{\dagger}$ 

 $^a$  Anisotropic thermal parameters are defined by:  $T = exp - (h^2 \beta_{11} + k^2 \beta_{22} + l^2 \beta_{33} + hk \beta_{12} + h l \beta_{13} + kt \beta_{23})$ .

		non-bonded distances in K <sub>2</sub> SbCl <sub>5</sub> <sup>a</sup>	
$Sb-Cl(1)$	2.799(2)	$Sb-CL(2)$	2.625(2)
$So-Cl(3)$	2.622(2)	$Sb-Cl(4)$	2.385(2)
$Sb-Cl(5)$	2.509(2)		
$C1(4)-Sb-C1(1)$	81.92(7)	$Cl(4)-Sb-Cl(2)$	83.54(8)
$CL(4)-Sb-CL(3)$	87.06(8)	$C1(4)-Sb-C1(5)$	87.77(8)
$CL(1)-Sb-CL(2)$	88.50(7)	$CL(2)-Sb-CL(5)$	90.43(8)
$C1(5) - Sb - C1(3)$	92.19(3)	$Cl(3)-Sb-Cl(1)$	87.22(8)
$C1(1)-Sb-C1(5)$	169.69(8)	$CL(2)-Sb-Cl(3)$	170.13(8)
$Sb-Sb$ '	3.932(1)	$Sb-CL(2)$ '	3.881(3)
$Cl(1)-Sb'$	3.701(2)	$CL(1)-CL(2)$	3.699(3)
$C1(2)-C1(5)i$	3.768(3)	$c1(1) - c1(3)$ <sub>ii</sub>	3.832(4)
$CL(1)-K(1)$ iii	3.187(3)	$CL(1)-K(2)$ iii	3.290(3)
$CL(2)-K(2)$ iii	3.168(3)	$CL(1)-K(1)iv$	3.247(3)
$CL(2)-K(1)iv$	3.147(3)	$Cl(3)-K(2)_{V}$	3.190(4)
$CL(2)-K(1)$	3.289(3)	$Cl(4)-K(1)$	3.317(3)

Table VIII. Bond lengths, bond angles, and significant

Primed atoms are those related by the center of inversion, as in Figure 4. Other symmetry operations referred to are: (i)  $-x$ ,  $\not\approx$   $y + y$ ,  $-z$ ; (ii)  $x$ ,  $y$ ,  $z-1$ ; (iii)  $x, \frac{1}{2} - y, \frac{1}{2} - 1;$  (iv)  $x, -\frac{1}{2} - y, \frac{1}{2} - 1;$  (v)  $1-x, -y, -z.$ 





Figure 4. Two inversion-related  $SbCl_5^=$  anions

 $\bullet$ 

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The antimony-chlorine bond distances in the base are all considerably longer than the axial distance. Two of these distances are almost identical (2.622, 2.625  $\stackrel{\circ}{\rm A}$ ) and involve chlorines  $t$ runs to each other. The other pair of chlorines form bonds of 2.799 and 2.509 A. length, with an average of 2.654  $\AA$ . In (NH<sub>4</sub>).<sup>SbCl</sup><sub>5</sub> all basul antimonychlorine bonds were found to be essentially of equal length (2.62  $\stackrel{\circ}{\mathtt{A}}$  average) while in  $[C^{\phantom{\dagger}}_{5}H^{\phantom{\dagger}}_{5}$ NH][ SbCl<sub>4</sub>] the corresponding bond length for chlorines trans to one another is  $2.63 \text{ Å}$ .

In the present study, the lengthening of the  $Sb-Cl(1)$ distance to 2.799 Å appears to be due to electrostatic repulsions and packing effects, as it is this chlorine that is involved in the closest non-bonded contacts; significant non-bonded distances are given in Table VIII. An antimony atom is only 3.701 Å from  $Cl(1)$ , whereas the sum of the van der Waal's radii is  $4.0$   $A.$  Also, the shortest chlorine-chlorine distance found,  $3.699$  Å, is between this chlorine and Cl(2)'. Both these distances would become significantly shorter if the Sb-Cl(1) bond were shortened. Thus, it can be inferred that this bond can be readily distorted if more efficient packing of the ions results. The shortening of the  $Sb-Cl(5)$  bond appears to be a truns effect in response to the lengthening of the  $Sb-CL(1)$  bond.

The observed bond lengths of the basal antimonychlorine bonds described above are consistent with characterizing the antimony bonding orbitals as being composed primarily of p-orbitals as in the three-center fourelectron bonding scheme described by Porter and Jacobson.  $52$ The lone pair appears to have a comparatively small stereochemical effect. This is supported by the relatively short antimony-antimony distance of  $3.932$  Å, where localized lone pairs, if present, would have to be directed toward one another. Distortions of the Cl-Sb-Gl bond angles from the expected 90° (cf. Table VIII) can be explained by interionic repulsions between inversion-related complexes.

The arrangement of the  $SbCl_{5}^{=}$  complexes in the unit cell is such that the chlorine atoms approximate a closest packed arrangement in the bc plane, as is evident in Figure 5, with the shortest chlorine-chlorine distances approximating the sum of the van der Waal's radii. None of the potassium ions lie between these planes, but rather are located among the  $SbCl<sub>5</sub>$  ions in the planes. The shortest potassium-chlorine distance is 3.147 Å.



# ALOP: A PROGRAMMED, SEMI-AUTOMATIC VERSION OP THE PSEUDO HEAVY ATOM TECHNIQUE

## Introduction

The increasing availability of high speed digital computers has enabled the crystallographer to make much more frequent use of mathematical techniques which previously were extremely time consuming. This fact has greatly enhanced the development and usage of direct methods, which recently have evolved to such an extent as to permit semi-automatic solution of crystal structures. In contrast, such indirect methods as Patterson superposition techniques have been employed more and more infrequently in recent times. Despite their great power, these latter methods require a great deal of human intervention and are somewhat tedious to perform. Clearly, what is required is a series of programs which would enable one to solve a structure using superposition techniques but requiring a minimum of human intervention.

This series of programs probably would have greatest applicability to, and thus should be specifically designed to solve, the pseudo heavy atom problem, in which the "heavy" atom does not contribute enough to the phasing to make heavy atom methods feasible. This problem occurs with compounds in which there are many light atoms and but

few heavy atoms, such as transition metal complexes or heavy atom derivatives of organic compounds.

In what follows, the essential features of this series of programs are detailed. The techniques employed naturally divide into two categories: (1) preparation of a Patterson map suitable for evaluation, and (2) utilization of this map to solve the structure.

## Map Preparation

A sharpened Patterson map<sup>11</sup> is created on either tape or disk by the program  $ALFF^{14}$  or any other similar Fourier series program. However, in this form the map is somewhat cumbersome to analyze. Hence, the program ALPP (Aines Laboratory Peak Picker) finds and outputs the heights and positions of all the peaks above a certain limit on the Patterson map. ALPP is divided into four subprocedures, the function of which is described below, and a main procedure, which controls the order of execution of the subprocedures. The program listing for ALPP is given in Appendix A.

Input to the first subprocedure, SETUP, consists of the size of the map and the minimum height of a peak which should be recognized. This subprocedure then reads and rewrites the map by layers, setting all points less than the minimum limit equal to zero.

The second subprocedure, LAYERS, reads these map layers and outputs layers with all non-peak grid points negated. The essence of this subprocedure is the subroutine PICKER, which is flowcharted in Figure 5. PICKER analyzes a line of the map point by point. The absolute value of each point is compared with that of the point before it. The lesser of the two is negated. LAYERS calls FICKER for each row and each column of the layer, and then compares points located diagonally to each point, negating those with smaller absolute values.

The next subprocedure, NTRLAY, compares the absolute value of a peak with those of the nine peaks adjacent to it in the next layer, again negating the smaller values. The only positive grid points which now remain are those of peaks. Next, these peak locations and heights are output and the modified map is printed, if desired for debugging purposes.

The final subprocedure, PKSORT, is a modified PL/I translation of QSTSRT. PKSORT sorts the peaks in order of decreasing height and outputs the ordered peaks.

The entire program runs in less core and requires somewhat less execution time than does ALPP. Thus, it is feasible to run the two programs back to back as one job.



Figure 5. Flowchart of subroutine PICKER

 $\bar{\beta}$ 

## Solution of the Structure

The main program of the series is ALOP (Ames Laboratory Origin Picker), and the program listing is given in Appendix B. ALOP is divided into a main procedure and seven subprocedures.

The main procedure reads in the size of the map, the number of peaks produced by ALPP, and the symmetry operations of the space group under consideration. These are in the form:  $p_x^x + q_x$ ,  $p_y^y + q_y$ ,  $p_z^z + q_z^2$ , where  $p_x$ ,  $p_y^2$ ,  $p_z = 1$  or  $-1$  and  $q_x$ ,  $q_y$ ,  $q_z = 0$  or  $\frac{1}{2}$ . Thus one can represent an equivalent position produced by: an inversion, a two-fold axis, a two-fold screw axis, a mirror plane in one of the three crystal directions or an  $\underline{a}$ -,  $\underline{b}$ -, or c-glide plane. Any of the triclinic, monoclinic, or orthorhombic space groups can be represented by these five symmetry operations (except Pdd2 or Fddd). In addition, any of the other 230 crystallographic space groups can be converted into a space group which only uses these operations, although some loss of symmetry generally occurs.

The first subprocedure, PART1, searches the peaks produced by ALPP for a suitable heavy atom peak. This peak must be among the largest 20% of the peaks. In addition, if any of its coordinates, u, v, w, is equal to 0 or  $%$  of the unit cell, then it must be at least twice as large as the largest peak which has no one of u, v, w

equal to 0 or  $\frac{1}{2}$ . These criteria assure that the heavy atom peak chosen is not solely a Harker vector, but allow the choice of a heavy atom which has 0 or  $%$  as one of its three coordinates, such as the antimony atom in  $K_2SbCl_{5}$ .

When PARTI finds a suitable heavy atom peak, it prints its coordinates,  $u_H$ ,  $v_H$ ,  $w_H$ , and uses it as the shift vector of an in-core algebraic superposition. For a peak u, v, w, to be retained, there must also be a peak at u' = u + u<sub>H</sub>,  $v'$  =  $v$  +  $v_H$ ,  $w'$  =  $w$  +  $w_H$  in the list of peaks produced by ALPP. In practice, since ALPP only determines the maximum grid point, and not the center, of each peak, these requirements are somewhat too stringent and thus u', v', w' is permitted to be off by one grid point in each direction. The values of  $u'$ ,  $v'$ , w' thus found are output for subsequent use.

PART2 assumes that the peaks produced by PART1 represent an electron density map of the compound, with its origin shifted to  $t^x, t^y, t^z, \underline{i.e.,}$  the following relations hold:  $u_i = x_i + t_x$ ,  $v_i = y_i + t$ ,  $w_i = z_i + t_z$ , where  $i = x_i + \tau_x, v_i = y_i + \tau_y, w_i = z_i + \tau_z,$  $x_j$ ,  $y_j$ ,  $z_j$  are the coordinates of the peak in electron density space. When the symmetry operations defined above are applied to a peak at  $x$ ,  $y$ ,  $z$  they produce an equivalent peak at  $p_x x + q_x$ ,  $p_y y + q_y$ ,  $p_z z + q_z$ . Thus for every peak at  $u = x + t_x$ ,  $v = y + t_y$ ,  $w = z + t_z$  there is an equivalent peak at u' =  $p_{x}x + t_{x} + q_{x}$ , v' =  $p_{y}y + t_{y}$ 

+  $q_y$ , w' =  $p_z^z$  +  $t_z$  +  $q_z^z$ . In Table X the results of subtracting and adding  $u, v, w$  and  $u', v', w'$  are given for all permutations of  $p_x$ ,  $p_y$ ,  $p_z$ . The asterisked (\*) sums determine values of  $t_x$ ,  $t_y$ , and  $t_z$ . PART2 and PART3 accumulate all possible values of  $t^x$ ,  $t^y$ , and  $t^z$  for the symmetry operations, as shown in Table XI. PART3 and PART4 write these values, along with the frequency of occurrence of each, on file SORTIN for future use. If any one of  $t_x$ ,  $t_y$ ,  $t_z$  is not determined, as is the case with all but inversion symmetry, it is set equal to -1. In addition to writing out the exact values of  $t^x, t^y,$  and  $t^z,$  those which are as much as one grid point off in each direction are also written out. Tables X and XI also demonstrate that values of  $t_x$ ,  $t_y$ ,  $t_z$  determined by two-folds, screws, mirrors, and glides have to meet criteria before they are acceptable, whereas inversions do not. Therefore, a greater weight is placed on these former values.

At this point, the following items may be present on file SORTIN:  $t^x, t^y, t^z$  triples;  $t^x, t^y, t^z, t^z$ ; or  $t^x$ ,  $t^z$  doubles; or  $t^x$ ,  $t^y$ , or  $t^z$  singles. These are sorted by the IBM procedure IHESRTA within their different types in order of decreasing frequency of occurrence. PART5 reads and stores 1/64 of the triples, 1/16 of the doubles, and 1/4 of the singles, all of which have the largest frequencies. These singles, doubles, and triples

	rable a. Relationships ased in determining $\alpha_{\mathbf{x}}, \alpha_{\mathbf{y}}, \alpha_{\mathbf{z}}$		
Quantity	$\mathfrak u$	$\boldsymbol{\mathrm{v}}$	W
$\mathbf U$	$A = x + t_x$	$B = y + t_y$	$C = z + t_z$
$U_{\rm O}$	$A + q_x$	$B + q_y$	$C + q$ <sub>z</sub>
$U_0 + U$	$2A + q_x$	$2B + q_y$	$2C + q_{z}$
$\sigma_{\rm O}$ – $\sigma$	$\mathbf{q}_{\mathbf{x}}$	$\sigma^{\rm A}$	$\mathtt{q}_{\mathtt{z}}$
$U_1$	$-x + t_x + q_x$	$B + q_y$	$C + q$ <sub>z</sub>
$U_1 + U$	* $2t_x + q_x$	$2B + q_y$	$2C + q$ <sub>z</sub>
$U_1 - U$	$-2x + q_x$	$\sigma^{\Delta}$	$\mathtt{q}_{\mathtt{z}}$
$U_2$	$A + q_x$	$-y + t_y + q_y$	$C + q$ <sub>z</sub>
$U_2 + U$	$2A + q_x$	* $2t_y + q_y$	$2C + q_z$
$U_2 - U$	$\mathbf{q}_{\mathbf{x}}$	$-2y + q_y$	$\mathtt{q}_{\mathtt{z}}$
$\mathbf{u}_3$	$-x + t_x + q_x$	$-y + t_y + q_y$	$C + q$ <sub>z</sub>
$U_5 + U$	* $2t_x + q_x$	* $2t_y + q_y$	$20 + q_2$
$U_3 - U$	$-2x + q_x$	$-2y + q_y$	$\mathtt{q}_{\mathtt{z}}$
${\tt U_4}$	$A + q_x$	$B + q_y$	$-z + t_{z} + q_{z}$
$U_4 + U$	$2A + q_x$	$2B + q_y$	* $2t_{z}$ + $q_{z}$
$U_4 - U$	$\mathbf{q}_{\mathbf{x}}$	$\mathrm{d}^{\pmb{\lambda}}$	$-2z + q$ <sub>z</sub>
$U_5$	$-x + t_x + q_x$	$B + q_y$	$-z + t_z + q_z$
$U_5 + U$	* $2t_x + q_x$	$2B + q_y$	* $2t_{z}$ + $q_{z}$
$U_5 - U$	$-2x + q_x$	$\sigma^{\text{A}}$	$-2z + q_z$

Table X. Relationships used in determining  $t$ ,  $t$ ,  $t$ 

Table X. (Continued)

 $\hat{\mathcal{A}}$ 



Type of symmetry element	Determines	Restrictions			
Plane perpendicular to x	$t_{\mathbf{x}}$	$v_1 - v = q_y, w_1 - w = q_z$			
Plane perpendicular to y	$t_y$	$u_2 - u = q_x$ , $w_2 - w = q_z$			
Plane perpendicular to z	$t_{\rm z}$	$u_{4} - u = q_{x}$ , $v_{4} - v = q_{y}$			
Axis in x	$t_y$ , $t_z$	$u_6 - u = q_x$			
Axis in y	$t_x$ , $t_z$	$v_{5} - v = q_{v}$			
Axis in z	$t_x$ , $t_y$	$W_2 - W = Q_2$			
Inversion	$t_x$ , $t_y$ , $t_z$	None			

Table XI. Restrictions placed on the determination of  $t_x$ ,  $t_y$ ,  $t_z$  by the type of symmetry element

are then merged with each other as described in Table XII, with a latitude of one grid point allowed in each direction. Thus, if a single in  $t_{\text{x}}$  equals the  $t_{\text{x}}$  of a  $t^x$ ,  $t^y$  double or is off by one grid point, then the frequency of the  $t^x$ ,  $t^y$  double is increased by the frequency of the single. A record is kept of each such merge.

This record is used in PART6, where all of the possible mergings for each symmetry element present are computed. Consider, as an example, the space group  $P2_{1}/c$  which has an inversion producing a triple  $t^x$ ,  $t^y$ ,  $t^z$ , a two-fold screw axis producing a double in  $t^x$ ,  $t^z$  and a c-glide producing a single in  $t_{\gamma}$ . PART5 should have merged a single in  $t_{\gamma}$  and a double in  $t_{\chi}$ ,  $t_{z}$  with each triple and left a record of having done so. The values of the triples which were not suitably merged are eliminated, and the triple remaining with the highest frequency of occurrence is then output for use in PART7.

Sometimes, however, a space group may not have an inversion. If it has three two-fold or screw axes, these will produce three doubles:  $t^{\text{xt}}_{x1}$ ,  $t^{\text{xt}}_{x2}$ ,  $t^{\text{xx}}_{x3}$ ,  $t^{\text{xx}}_{x3}$ , In the cases where  $t^{\text{}}_{x1} = t^{\text{}}_{x2}$ ,  $t^{\text{}}_{y1} = t^{\text{}}_{y3}$ , and  $t^{\text{}}_{z2} = t^{\text{}}_{z3}$ , a triple is created with a frequency equal to the sum of the frequencies of the three doubles. The highest such triple is then taken as the origin. If any one of  $t_{\mathbf{x}},$  $t_y$ , or  $t_z$  is not determined by the above procedures, it



 $\sim$ 

Table ZII. Merging of singles and doubles into doubles and triples

is set to zero, as the origin may have any value in that direction.

PART7 computes the electron density space coordinates of all the peaks found in PARTI by using the equations:  $x_i = u_i - t_x$ ,  $y_i = v_i - t_y$ ,  $z_i = w_i - t_z$ . For each peak, this procedure then attempts to find the symmetry equivalent peaks. If at least half of the equivalent peaks are present, a peak is output with the appropriate fractional coordinates.

ALOP uses less than 32K words of main core storage and runs in less than five minutes of CPU time for a 32 **X** \$2 **X** 52 map with less than 500 peaks input from ALPP.

#### Results

The programs ALPP and ALOP were tried on two different problems and the technique was successful in both cases. In the first case, the structure of  $D-(1,5)$ -glucono lactone (space group  $P2_1^2^2^1$ )<sup>33</sup> was "chlorinated" by substituting chlorines for two of the hydrogens, yielding a molecule with two chlorines, six oxygens, and six carbons. The heavy atom ratio for this compound was  $\sum_{\text{heavy}}^{2}$  / $\sum_{\text{light}}$ )<sup>2</sup> = 0.963. All of the non-hydrogen atoms were used in a structure factor calculation and a sharpened Patterson map was computed from these structure factors. ALPP found 502 off-origin peaks in this map with heights rang

ing from 13 to 234. A heavy atom peak, which, proved to be an oxygen-chlorine vector, was found at  $u = 30$ ,  $v = 26$ ,  $w = 3$  with a height of 99 on the  $32 \times 32 \times 32$  map. In-core superposition retained 187 peaks with a minimum height of 23. This is a good demonstration of how the superposition process tends to eliminate spurious peaks. Multiple merging produced eleven possible origins, of which the one with the highest frequency of occurrence was  $t_x = 14$ ,  $t_{y}$  = 13,  $t_{z}$  = 12. When this value was used as an origin, 33 symmetry independent peaks were retained. Subsequent least square refinement of the positional parameters and atom multipliers of these 33 positions through twelve least squares cycles resulted in a conventional discrepancy factor of  $R = 0.206$ . The scattering factor curve for this and the subsequent test case is defined by:

$$
\hat{\mathbf{f}}_{\hat{\mathbf{i}}} = \sum_{\hat{\mathbf{j}}=1}^{N} \mathbf{f}_{\hat{\mathbf{i}}, \hat{\mathbf{j}}} / \sum_{\hat{\mathbf{j}}=1}^{N} z_{\hat{\mathbf{j}}},
$$

where  $f_{i,j}$  is the normal scattering factor for the j<sup>th</sup> atom over a range of sin $\theta/\lambda$ ,  $^{15}$   $Z_{\tilde{J}}$  is the atomic number of the atom, and the sum is over all the atoms in the unit cell.

The second test case was that of  $K_2SbCl_5$  which crystallizes in space group  $P2_{1}/c$  with four eight-atom moieties per unit cell. The heavy atom ratio for this compound is 1.200. ALPP found 258 off-origin peaks on the sharpened

Patterson map, ranging in height from 17 to 508. A heavy atom peak, which was produced by an antimony-antimony vector, was found at  $u = 20$ ,  $v = 16$ ,  $w = 3$  with a height of 282 on the 32 x 32 x 32 map. In-core superposition retained 112 peaks. The origin with the highest frequency of occurrence was at  $t_x = 10$ ,  $t_y = 0$ ,  $t_z = 9$ , and its use yielded 20 symmetry independent peaks. Sixteen cycles of full matrix least squares refinement of the positional parameters and atom multipliers of these positions led to the removal of eleven of them as their atom multipliers dropped to less than 0.4-0. The remaining nine positions refined to yield a conventional discrepancy factor of **R = 0.240.** 

PSST: A COMBINATION OF PATTERSON SUPERPOSITION AND SIGMA-2 TECHNIQUES FOR PHASE DETERMINATION

Introduction

At present there are two methods commonly used for solving structures of moderate complexity. Direct methods, such as those using the  $\Sigma_2$  relation,  $^{34}$  are primarily reciprocal space methods, while methods based on the deconvolution of the Patterson function (for a detailed bibliography see Buerger $^{55}$ ) are primarily real space methods.

There is a certain parallelism between these two methods and both have their advantages and disadvantages. Both involve some initial choices which can greatly influence the success of the method. The  $\Sigma_2$  relation can be readily programmed and many automatic or semi-automatic procedures have been developed based on this relation. However, it is usually difficult to make use of known structural information in this approach. Furthermore, if no reasonable structure is produced, the investigator has little recourse but to modify his choice of origindetermining reflections or to closely examine those phases selected at the early and most critial part of this phase determination procedure. With Patterson methods, on the other hand, there are usually a greater number of options initially available corresponding to differing modes of

selection of peaks for superposition. Also, known structural information can be more readily introduced, since a real space representation is being employed. For complex structures, however, large numbers of superpositions are generally required and severe degradation of the structural image can result due to the accumulation of errors in the atom position selections.

The advent of the fast Fourier transform technique and its application to crystallography<sup>36</sup> allows rapid real space - reciprocal space conversions. Thus, it is appropriate to consider whether a hybrid direct method - Patterson method approach having greater power than either method alone can be developed. It is just such a hybrid that is described here.

#### Theoretical Basis of the Method

In direct methods employing the  $\Sigma_2$  relation, one starts with E's of large magnitude so that phases will be determined with reasonably high probabilities. Due to the fact that the initial contributors are few in number, erroneous phase indications may still result which then propagate in the phase determination procedure. Consider the map produced by a Fourier transform using as non-zero coefficients only those reflections input into the  $\sum_{2}$  relation (henceforth this map will be referred to as the  $\Sigma$  g map). For the phase determination procedure to be successful, it

is necessary that the square of this  $\Sigma_2$ -map better resemble the true electron density function than does the  $\Sigma_2$ -map itself. If so, the phases of the transformed coefficients with larger magnitudes should be good approximations to the phases of the corresponding large |E|'s.

Since the initial set of E's is very small in number, possibly only the three origin-determining reflections, it is the characteristics of the  $\Sigma_2$ -map which should be closely examined. First, if large |E|'s are used, there is a high probability that atoms will lie on, or at least near, some of the maxima occurring on this map. Second, there will be a large number of extraneous maxima on the square of the  $\Sigma_2$ -map -- a number far exceeding the number of atoms in the cell. The first is an asset and the second a liability for successful phase determination. These extraneous peaks come from either maxima or minima on the  $\Sigma_2$ -map. It is the presence of these extra peaks that greatly inhibits the production of additional phase information.

One would like to modify the square of the  $\Sigma_0$ -map so that it better resembles the electron density function. An obvious such modification is to eliminate all negative regions (or possibly regions less than a certain lower bound) by setting these regions equal to zero before squaring. Indeed, this technique has been applied by Barrett

37 and Zwick<sup>27</sup> to some extent in the extension and refinement **of crystallographic protein phases.** 

In the initial application of the  $\Sigma$ <sub>2</sub> relation, there **are so few phased E's that even with the modification described above, there are large positive regions in many areas of the map which do not correspond to atomic locations. However, the Patterson function contains images of the structure with each atom in turn at the origin. Therefore, if an atomic position can be selected on the ^-map, a superposition of the appropriately scaled Patterson map on this position should produce a map with considerably fewer extraneous peaks and with much better resemblance to the electron density function. If such a map is then squared to emphasize the higher peaks and transformed, better and more extensive phase information will result. Once a sufficiently large set of phased S's has been obtained, the remaining |E|'s can be efficiently phased by normal application of phase extension and refinement** techniques of the  $\Sigma$   $\rightarrow$  type.

## **Application of the Method**

#### **Centric case**

The compound  $\beta$ -picoline-N-oxide fumaric acid adduct<sup>38</sup> **was chosen as a centrosymmetric test case. We felt this would be a good test case as the structure had previously been solved only with difficulty by use of a roving model —** 

vector verification technique after usual superposition and symbolic addition procedures had not proved successful.

This adduct crystallizes in the space group  $P2<sub>1</sub>/c$ with  $\underline{a} = 3.888$ ,  $\underline{b} = 14.194$ ,  $\underline{c} = 14.666$   $\underline{a}$  and  $\beta = 98.85^\circ$ . There are two adducts of formula  $C_{16}H_{18}N_2O_{12}$  per unit cell, i.e., the adduct has a center of symmetry. Values of  $|E|$ were computed in the ordinary way and there were 270  $|E|$  's greater than 1.0. The three selected origin-determining reflections are given in Table XIII, These three, along with the symmetry-related reflections, were used in a  $\Sigma$ <sub>2</sub>-map calculation (32 x 52 x 32 grid points). Some typical sections of this map are shown in Figure 7.

The positions of the five highest peaks on this map were noted and in each case checked by examining the Patterson for the presence of vectors between symmetryrelated atoms. In each case, such vectors were found and indeed these peaks fell approximately (within  $0.5 \n0$ ) at the positions of four carbons and a nitrogen in the structure.

All of the negative regions in the  $\Sigma_2$ -map were then set equal to zero and five independent superpositions were carried out by placing the origin of the Patterson at each of the five peak positions and the symmetry-related points. A sharpened Patterson function was used and the minimum procedure  $emploped.$ <sup>39</sup> The scale of the Patterson was







Figure 7a. Composite of layers 3.4,5 of a  $\Sigma_{2}$ -map for  $P$ -picoline-N-oxide fumaric acid adduct. contoured at 40 and 80. ---- contoured at  $-40$  and  $-80$ . Atom positions indicated by X.



Figure 7b. Layer 9 of a  $\Sigma_2$ -map for  $\beta$  -picoline-N-oxide fumaric acid adduct. - contoured at 15 and  $25.$  ---- contoured at  $-15$  and  $-25.$  No atoms in or near this section.

adjusted so that the value of the highest off-origin peak on the Patterson map was approximately equal to the value of the highest off-origin peak on the  $\Sigma_{\supset}$ -map. Figure 8 shows some typical results using the peak corresponding to the nitrogen atom position as the superposition point.

Each of the five maps was then squared and the resulting transform coefficients  $(G<sub>hk1</sub>)$  calculated using a fast Fourier algorithm. $14$  Only the 50 reflections with the largest values of the product  $|E*G|$  and with  $|E| > 1.50$ were kept. Some typical results are shown in Table XIV. The signs determined for each hkl were then averaged. (In some cases more than 80% of the 50 largest |E]'**S** were signed correctly in the individual sets.) These signed reflections were then used to compute a new  $\Sigma_2$ -type map, and negative regions were set equal to zero. The transform of the square of this map was used to obtain signs (92% of which were correct) for input at the 90% confidence level into a phase extending program.<sup>40</sup> The resulting complete set of 270 phases was used to compute a map from which the positions of all the non-hydrogen atoms in the structure were readily determined. There were no significant spurious peaks on this map. This method has recently been used to solve the structure of diisopropyl- (2,3,4,5-tetraphenyl-cyclopenta-2,4-diene)-phosphate, / which crystallizes in space group  $P2<sub>1</sub>/c$  with 160 non-

$\mathbf h$	k	$\mathbf{1}$	E	ı	$\mathsf{S}$	$\overline{3}$	Calculated signs 4	5	Known sign
$-2$	10	4	3.25		$\div$	$\ddot{}$	$\div$	$\begin{array}{c} + \end{array}$	$\div$
$-1$	$\mathbf{L}$	$\overline{7}$	3.13		$\bf +$	$+$		$+$	$\ddot{}$
$\circ$	$11\,$	9	3.00			$\ddot{}$			$\div$
$\circ$	12	4	2.94				$\div$		
$-1$	10	12	2.80						
$\mathbf 1$	$10$	$\rm ^8$	2.73	$\ddot{}$					
$-2$	4	12	2.63				$\div$		$+$
$\mathbf 1$	$11\,$	$\overline{7}$	2.55	$+$					
						$\mathbf{r}$			

Table XIV. Representative results<sup>a</sup> obtained for a centric test case ( $\beta$ -picoline-N-oxide fumaric acid adduct)

The absence of a calculated sign indicates that the value of  $|E^*G|$  for that reflection was smaller than the limit used.



Figure 8a, Results of Patterson map superimposed on the  $\Sigma$ <sub>2</sub>-map of Figure 7 at the four nitrogen equivalent positions. Composite of layers 3,4,5. contoured at 9 and 18. Atom positions indicated by X.



Figure 8b. Results of Patterson map superimposed on the  $\Sigma$ -map of Figure 7 at the four nitrogen equivalent positions. Layer 9. - contoured at 5. No atoms in or near this section.

hydrogen atoms per cell. $41$ 

## Acentric case

The compound chosen as a test for the acentric case was D- $(1,5)$ -glucono lactone,<sup>33</sup> which crystallizes in the space group  $P2^12^1$ , with axial lengths  $a = 7.838$ ,  $b = 12.322$ ,  $c = 7.544$   $\text{\AA}$ . Of the 1259 unique observed reflections, 435 have  $|E|$  greater than 1.0. Three origin-determining reflections were selected ( $|E| = 2.39$ , 2.51, 2.04), two of which were non-zonal reflections. To make the test a reasonably severe one, no enantiomorph-selecting reflection was included; the enantiomorph was chosen as part of the superposition procedure.

The  $\Sigma$  <sub>2</sub>-map produced from these and their symmetryrelated reflections showed more than 25 large peaks, all of which had relative heights between 120 and 185. The relative height of the other peaks on this map was less than 90. However, only eight of these 25 peaks generated appropriate Harker vectors on the Patterson. Of these eight, four were clearly in one enantiomorph, since they remained when the superposition point was a member of the set.

Using these four peaks and proceeding in a manner similar to the centric case above, the average phases of the 30 largest  $|E^*G|$ 's with  $|E| > 1.40$  were obtained. Some typical results for this acentric structure are given in

**64-**

Table XV. The transform of the square of the expanded  $\Sigma$ <sub>2</sub>-map yielded initial phases for the 50 largest  $|E*G|$  's with  $|E| > 1.40$ . These phased E's were then input into a phase-extending program at the 90% confidence level. In the resulting set of 410 phased E's, 390 were within 40° of the published phases and their transform readily revealed all the non-hydrogen atomic positions. As in the centric case, there were no significant spurious peaks on this final map.

## Extensions of the Method

Since a reasonably severe test of the method was desired, only three origin-determining reflections were used for the initial  $\Sigma_{\rho}$ -map calculation in each case. Other reflections, phased by either  $\Sigma_2$  or  $\Sigma_1$  relations (and perhaps with appropriate weights), can also be included to produce peaks of greater reliability. In the centric case, those  $\Sigma_1$  reflections which were determined with greater than 90% probability were included in the calculation of the  $\Sigma_2$ -map. Their addition improved the phase agreement to a small extent. The frequency check procedure<sup>42</sup> can also be applied to the  $\Sigma_2$ -map.

It is also possible to include an additional strong reflection in the origin-determining reflections, thus producing two  $\Sigma_{\supset}$ -maps in the centric case, for example. The resulting two sets of averaged phases could then be




The absence of a calculated phase indicates that the value of IE\***G**I for that reflection was smaller than the limit used.

handled in a similar way as in those direct methods programs which produce multiple solutions.

Lower bounds different from zero can be used for modification of the resultant superposition map. Indeed, some calculations for the centric example have shown that use of a lower bound of 10% of the maximum peak height results in somewhat improved agreement.

From Tables XIV and XV, it is obvious that a straight average is a poor procedure compared to a more judicious method of averaging. Preliminary computations indicate that in the centric case a requirement of a net of two like signs produces a significant improvement in the percentage of correct signs selected.

The phases are also somewhat sensitive to grid point resolution of the maps. In the centric structure, calculations were run using a grid of 0.12 x 0.45 x 0.45 Å resolution. Results improve if a resolution of approximately  $0.25$   $\stackrel{\circ}{\text{A}}$  is used in all three directions.

Squaring the resultant superposition map appears to give the best phases compared with other powers which could be employed. All exponents from 0.2 to 6.6 in steps of 0.2 were examined for a typical run, with 2.0 appearing the most satisfactory.

## RESEARCH PROPOSALS

The proposals for research given below suggested themselves during the course of the research described in this thesis, No claim of an exhaustive literature search for these ideas is implied, although it is the author's belief that they are not presently under investigation.

Due to the fact that more investigators without formal training are employing crystallography as a research tool, there is a desperate need for the development of computer programs which can make many of the routine decisions now requiring human intervention. ALPP eliminates, or at least makes much less tedious, the contouring of a map, and is an example of this type of program. Such a program should be developed to implement the PSST method. One could write a program to implement the heavy atom technique, which would work in conjunction with ALOP, by cycling structure factor and Fourier series calculations, at the same time deciding whether or not possible atomic positions should be retained. As well as being easy for the initiate to use, the above programs should also permit enough intervention so that the course of their execution can be altered by the more experienced user.

The prime difficulties in implementing direct methods are in choosing between various sets of phases

produced and in discriminating spurious from real atomic peaks. The first problem is being extensively researched and such methods as the comparison of absolute figures of merit and the psi zero test of MULTAN,  $40$  the absolute value of rho test reported by W. Ozbirn,  $43$  and a plot of psi zero vs. the absolute figure of merit as described by Cochran and Douglas $44$  are all examples of possible solutions. A promising approach to the problem of peak discrimination should be the superposition of the Patterson on a likely atomic position, similar to the technique used in the PSST method. Another approach would be the development of a program which makes use of known interatomic vectors, such as those within a benzene ring, to choose peaks of chemical sense. A third would be the development of a more exact and easily used technique for producing three-dimensional models. A model constructed of wooden spheres suspended by fine lines from a perforated plate would satisfy these requirements.

The relative ease in determining the bonding mode of cyanide in  $trans-(CN)$ <sub>2</sub> trien cobalt (III) perchlorate indicates that X-ray diffraction data may be sensitive enough to distinguish between carbon and nitrogen. A more intensive effort on the part of investigators on this problem and perhaps even a reexamination of published data could be initiated.

The number of antimony (III) structures which have been determined is exceedingly small. The investigation of a number of different compounds of formula  $M_2Sb^{III}X_5$  $(M = Na, K, Rb, Cs; X = Cl, Br, I)$  would be of interest in determining the influence of the  $M^+$  and the  $X^-$  on the configuration of the  $SbX^{\equiv}_{5}$  complex.

The  $\Sigma_{\supset}$ -map has a much wider applicability than is described in the PSST method. In combination with the discriminator function $^{45}$  to test possible atomic positions, it should have great power. In addition, using the expanded starting set of phased E's to produce a map which then would be used in place of the  $\Sigma_{\rho}$ -map, is an extension which has promise.

APPENDIX A

PROGRAM LISTING OP ALPP

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 $\label{eq:2.1} \frac{1}{\sqrt{2}}\int_{\mathbb{R}^3}\frac{1}{\sqrt{2}}\left(\frac{1}{\sqrt{2}}\int_{\mathbb{R}^3}\frac{1}{\sqrt{2}}\left(\frac{1}{\sqrt{2}}\int_{\mathbb{R}^3}\frac{1}{\sqrt{2}}\right)\frac{1}{\sqrt{2}}\right)\frac{1}{\sqrt{2}}\,d\mu.$ 

 $\mathcal{Z}$ 



\* PROCESS('ATR,XREF,NOSTMT,OPT=02');<br>/\* CREATES A TIME - DATE STAMP FOR FUTURE USE<br>(SUBRG,STRG): TM DT: PROCE<u>DURE;</u> **COBRUS SIRUMP CHARACTER(54) EXTERNAL,**<br>DECLARE MOS(12) CHARACTER(9) VARYING **\*/ TMDT0020 DECLARE DECLARE •MARCH\*,'APRIL' ,"JUNE','JULY' 'OCTOBER','NOVEMBER','DECEMBER'');<br>TM CHARACTER(9), TMS(2) CHARACTER<br>DT CHARACT<u>ER(6), DS</u> (3<u>)</u> CHARACTER! ,'MAY'**  AMPM CHARACTER(4);<br>INITIAL ('JANUARY<u>','FEBRUAR</u>Y', **DECLARE TM CHARACTERIOI, TMS(2) CHARA**<br>DECLARE DT CHARACTER(6), DS (3) CHAR<br>IF (TMS(1) > 11) THEN AMPM = "P.M." **CHARACTBRI2) CHARACTER!2) 'AUGUST','SEPTEMBER' , DEFINED TM; TM = TIME; DEFINED DT; DT = DATE; ELSE AMPM = 'A.M.'; TMDTOOIO TMDT0030 TMDT0040 TMOT0050 TMDT0060 TMDT0070 TMDT0080 TMDT0090 TMDTOIOO** 



STAMP = 'RUN ON ' || MOS(DS(2)) || ' || DS(3) || ', 19' || DS(1) || TMDTO110

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APPENDIX B

PROGRAM LISTING OF ALOP

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ALP00460 **STAMP** CHARACTER (54) EXTERNAL; **DECLARE DECLARE** SYMMPO (#SYMM) CONTROLLED EXTERNAL, AL PO0470  $\mathbf{1}$  $2$  (PX, QX, PY, QY, PZ, QZ) AL PO0480 FIXED BINARY (15); (#PĖAKŠ,#ŠYMM,#ŔTÑĎ,IACR,IDWN,ILAY)<br>FIXED BINARY (15); ALP00490 **DECLARE ALP00500 ALP00510** BIT (1) EXTERNAL: (SGNL, NCNT) **DECLARE ALPO0520** SGNL='1'B IF THERE IS A SYMMETRY ELEMENT WITH PX=PY=P2=-1.<br>NCNT='1'B IF THERE ARE SYMMETRY ELEMENTS WITH ONLY: PX=-1, **ALP00530 ALP00540**  $PY=-1, PZ=-1, PX=PY=-1, PY=PZ=-1, PZ=PX=-1.$ AL PO0550 \*/ ALP00560 FIXED BINARY (31): AL PN0570 **DECLARE** (ITXYZ,LNG, IERR)  $(\tilde{C}H\tilde{A}\tilde{R}U\tilde{5}\tilde{3})$ ,  $\tilde{C}H\tilde{A}\tilde{R}\tilde{U}\tilde{3}\tilde{9})$ , **ALPO0580 IHESRTA ENTRY DECLARE** FIXED BIN(31), FIXED BIN(31)); ALPO0590 CHARACTER (53); ALPN0600 **SFLD DECLARE** ALP00610 **RFLD** FILE RECORD SEQUENTIAL; **ALP00620 DECLARE** SORTIN CHARACTER (100) VARYING: ALP00630 **DECLARE SDBG** ALP00640 EXTERNAL BIT (1); **DECLARE** DBG. BIT (1) EXTERNAL: ALP00650 **DECLARE DNV**  $EISE$  DNV = '0'B; **ALPN0660** IF (LENGTH(SDBG)  $> 1$ )  $ELSE$  DBG = '1'B; ALPO0670  $(LENGTH(SDBG) = 0)$ THEN DBG =  $'0'$ Bi 1 F **ALPO0680** CALL TM DT: ALP00690  $LNG = 24000$ **ALPO0700**  $\overline{RFL}$ D='  $\overline{REC}$ ORD TYPE=V, LENGTH=(24, 24, 24, 24, 24) EXIT: ALP00710 ON SUBSCRIPTRANGE SNAP BEGIN; PUT DATA (I); **END: ALPO0720 ALP00730** READ AND INTERPRET THE CONTROL CARDS \*/ ALPO0740 THEN PUT EDIT ('DEBUGGING OUTPUT WILL BE PRINTED.') **ALPO0750** IF (DBG) **ALP00760** (SKIP(4),A); ELSE; **ALP00770** (TITLE, #PEAKS, IACR, IDWN, ILAY, #SYMM) GET EDIT **ALP00780**  $A(80), 5$   $F(4)$ : PUT EDIT (TITLE, STAMP, ALP00790 INUMBÉR OF POINTS ACROSS A FULL LINE =', IACR, **ALPO0800 ALP00810** INUMBER OF LAYERS IN A FULL MAP = ,ILAY, ALP00820 \*NUMBER OF PEAKS IN INPUT DATA SET = \*\* \* PEAKS. ALP00830 "NUMBER OF SYMMETRY CARDS TO BE READ =", #SYMM) **ALPO0840**  $($ A(80), SKIP(3), A(54), 5 (SKIP(3), X(20), A, F(4))); **ALP00850 ALPO0860** ALLOCATE SYMMPO;<br>DO I = 1 TO #SYMM BY 1; ALP00870  $(PX(1), QX(1), PY(1), QY(1), PZ(1), QZ(1))$ **ALPO0880** GET EDIT **ALP00890**  $SKIP, 6$   $F(4)$  : PUT EDIT (ISVMMETRY OPERATION, I, PX(I), \*\*X +\*, QX(I), '/2';<br>PY(I), \*\*Y +\*, QY(I), '/2', PZ(I), \*\*Z +\*, QZ(I), '/2';<br>SKIP(2), X(25), A, F(3), 3 (F(7), A, F(2), A)); **ALP00900** AI PN0910 **ALP00920 ALP00930** END SLUP: ALP00940 GET SKIP;

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SLUP:

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**ALP00950** 



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 $\bar{\mathcal{A}}$ 







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JLUP:

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\* PROCESS ( ' ATR, XREF, NOSTMT, OPT=02' 1;

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 $\mathcal{L}(\mathcal{L}(\mathcal{L}))$  and  $\mathcal{L}(\mathcal{L}(\mathcal{L}))$  . The contribution of  $\mathcal{L}(\mathcal{L})$ 

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 $\mathcal{L}(\mathcal{L}^{\mathcal{L}})$  and  $\mathcal{L}^{\mathcal{L}}$  are the set of the set of  $\mathcal{L}^{\mathcal{L}}$ 





 $-P2-3520$ <br>-P2-3530

RETURN;<br>END PART2;



TOT


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WRITTEN QUT'ON'THE FILE "SORTIN", WITH THE UNDETERMINED VALUES-P4-0090<br>SET EQUAL TO -1.





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 $\mathcal{A}^{\mathcal{A}}$ 







 $\frac{1}{2}$ 

RRD:  $\rightarrow$ 

LOT



 $\mathcal{L}^{\text{max}}_{\text{max}}$  .



 $\begin{bmatrix} \n\text{DY} & = & \text{ABS} & \text{I} & \text{SY} & - & \text{SY} & \text{J} \\
\text{IDZ} & = & \text{ABS} & \text{I} & \text{SZ} & - & \text{SZ} & \text{J} & \text{J} \\
\end{bmatrix}$ **101 = IOY; 102 = IDZ; IF (15=31 THEN IF (IDY>1)**   $\tilde{B}$ <sup>F</sup>((IDY<2)<sup>-</sup> $\tilde{B}$ (IDZ<2))<sup>-+</sup>1'B; **IF (IS=5) BYZ( THEN IF FIDZ>L) BYZ( J)=\*L\*; IF ((IS=7)6(I0Y<2)6( IDZ<2)» CMPXYZ: IF (ID1=0) ELSE IF**  (102=0) THEN IF (102=0) **0K(J) = OK(J) + (P » FREQI \* FREQ(J)) CK(7): NOJ: END ILUP;<br>
END ILUP;<br>
OK = OK \* ((1000.000 \* FI<br>
IF (DBG) THEN DO;<br>
PUT EDIT ('SAVED TX, TY,<br>
PUT EDIT (' <u>TX TY TZ</u> SY' » TX TY TZ SY MAPPING»)(SKIP(3), FREQL TZ AFTER FREQ TX TY TZ SY FREQ TX TY TZ SY FREQ TX TY TZ SY FREQ', FREQ»)**   $\begin{array}{cc} (SKI\overrightarrow{P}(2), \overrightarrow{A}_1\overrightarrow{A}_1\overrightarrow{A}) \\ (RESULT \overleftrightarrow{I}_1\overleftrightarrow{O}_0)^{\frac{1}{2}} = 1 \end{array}$ **FREE STORE;**  END PARTS; **ELSE DO: THEN ELSE 00; THEN GO TO GO TO CMPZ; CMPXYZ; ELSE GO TO NDJ;**   $\frac{1}{2}$ **HEN**  $P = PQ$ ; **ELSE THEN ELSE p = gn**  THEN GO TO NDJ; **DO; GO TO GO TO CMPY; NDJ;<br>
BYZ(I) = '1'B;**<br>
BYZ(I) = '1'B; **NDJ; BYZ(I) = •1\*B; END; END; ELSE;**   $P = P1;$  $p = p2;$  $\overline{0}K = \overline{0}K$  **\*** ((1000.000 \* FLOAT(#XYZ)) / SUM(OK)); **FRED TX TY TZ SY**<br>PUT SKIP(2);<br>I = 1 TO #XYZ))(4 F(3),6 B(1) **END;**  -P5-1840<br>
-P5-1850<br>
ELSE; -P5-1860<br>
-P5-1870<br>
-P5-1880<br>
-P5-1900<br>
-P5-1900 P5-1900<br>-P5-1910<br>-P5-1920<br>-P5-1930 **END; ELSE; -P5-1940<br>P5-1950 -P5-1950**<br>P5-1970 -P5-1970 -P5-1980<br>-P5-1990<br>-P5-2000<br>-P5-2020<br>-P5-2030<br>-P5-2040 -P5-2040<br>
-P5-2050<br>
-P5-2060<br>
-P5-2070<br>
-P5-2080<br>
-P5-2110<br>
-P5-2120<br>
-P5-2120<br>
-P5-2130<br>
-P5-2130 **′−P5−2120<br>P5−2130**<br>P5−2140 **-P5-2150 A)** 

-P6-0020<br>-P6-0020<br>-P6-0030<br>-P6-0060 -P6-0060<br>-P6-0080 -P6-0080<br>-P6-0110<br>-P6-0120<br>-P6-0130<br>-P6-0140 \* PROCESS('ATR,XREF,NOSTMT,OPT=02');<br>/\* CHOOSE THE MOST FREQUENT VALUE OF TX, TY, AND TZ \*/ -P6-0020 **PART6: PROCEDURE (IACR, IDWN, ILAY);<br>/\*** THIS SUBROUTINE CHECKS THROUGH ALL OF THE PEAKS MAPPED INTO EACH OTHER TO FIND THE HIGHEST SYMMETRY ELEMENT INTO WHICH ALL THE UTHER TO FIND THE FIGHTS HAVE BEEN MAPPED, WHICH HIGHER SYMMETRY<br>ELEMENT ALSO HAS THE HIGHEST FREQUENCY. IF THIS SYMMETRY ELEMENT<br>IS SUCH THAT ALL OF THE OTHER SYMMETRY ELEMENT TYPES HAVE BEEN. MAPPED INTO IT, THE PROGRAM SETS THE APPROPEIATE VALUES OF TX, IX, MAPPED INTO ITS THE PROGRAM SETS ANY ONE OF TX, TY, OR TZ UNDETER-<br>AINED, IT IS SET TO ZERO, AS THE ORIGIN IN THAT DIRECTION MAY BE<br>SET TO ANY VALUE. IF THE SYMMETRY ELEMENTS DETERMINE ALL THREE



TTT

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 $\begin{array}{c}\n \cdot \text{r} \\
 \uparrow \text{r} \\
 \end{array}$ 



**NDJL:** 



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 $\frac{1}{4}$ 

 $\mathcal{F}_{\mathcal{F}}$ 113









**-P6-3150** 









 $\mathcal{L}(\mathcal{L}(\mathcal{L}))$  and  $\mathcal{L}(\mathcal{L}(\mathcal{L}))$  . The contribution of  $\mathcal{L}(\mathcal{L})$ 



 $\label{eq:2.1} \frac{1}{2} \sum_{i=1}^n \frac{1}{2} \sum_{j=1}^n \frac{$ 



 $\mathcal{L}(\mathcal{L})$  and  $\mathcal{L}(\mathcal{L})$  . The set of the set of  $\mathcal{L}(\mathcal{L})$ 

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